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Gases Class SHORT ANSWER Answer the following questions in the space provided.

b Pressure — orce For a constant force, when the surface area is tripled the surface area pressure is (a) doubled. as much. (c riplied. 7-0 (d) unchanged. Rank the following pressures in increasing order. (c) 76 torr (a) 50 kPa O, OOìctbv-x

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SHORT ANSWER Answer the following questions in the space provided. 1. The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided

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by 22.4 L. 2.

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Answers**

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ANSWER . Answer the following
questions in the space provided. 1. c The
molar mass of a gas at STP is the density
of that gas (a) multiplied by the mass of

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1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided by 22.4 L. 2. c For the expression $V = nT$,

CHAPTER REVIEW Gases

462 Chapter 11 Gases Discovering the Relationships Between Properties If we want to explain why a weather balloon

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carrying instruments into the upper atmosphere expands as it rises, we need to consider changes in the properties of the gases (pressure, volume, temperature, or number of gas particles) inside and outside the balloon.

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CHAPTER 11 REVIEW. Molecular
Composition of Gases. MIXED REVIEW.
SHORT ANSWER Answer the following
questions in the space provided. 1. cThe
average speed of a gas molecule is most
directly related to the . (a)polarity of the
molecule. (b)pressure of the gas.
(c)temperature of the gas. (d)number of
moles in the sample.

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Answers Section Goals and Introductions

Section 111 Gases and Their Properties

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Goals To describe the particle nature of both real and ideal gases To describe the properties of gases that can be used to explain their characteristics: volume, number of particles, ...

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(25) definition & applications of pressure (also atmospheric) SI unit of force.
definition & use of a barometer.
standard temperature & pressure (STP)
Definition of Dalton's law of partial pressures. Definitions & formulas for Boyle's, Charles', Gay-Lussac's, and combined gas laws

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CHAPTER 11 REVIEW Gases SECTION 2
SHORT ANSWER Answer the following questions in the space provided. 1. State

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whether the pressure of a fixed mass of gas will increase, decrease, or stay the same in the following circumstances:

increase a. temperature increases, volume stays the same decrease b. volume increases, temperature stays the same

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School

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11 REVIEW Gases SECTION 3 SHORT
ANSWER Answer the following questions
in the space provided. 1. _____ The molar
mass of a gas at STP is the density of
that gas (a) multiplied by the mass of 1
mol. (c) multiplied by 22.4 L. (b) divided
by the mass of 1 mol. (d) divided by 22.4

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L. 2. _____ For the expression , P

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Chapter 11 - Gases Chapter 11 focuses on gas behavior and the gas laws. In Chapter 10, students were given an overview of the kinetic-molecular theory of matter and discussed how this

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have not yet neon gas mixed with argon gas (Objs 21 & 22) a If the total pressure of the mixture of gases is 130 kPa and the partial pressure of neon

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CHAPTER 11 REVIEW. Gases. SECTION 3.
SHORT ANSWER Answer the following

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questions in the space provided. 1. The molar mass of a gas at STP is the density of that gas. (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided by 22.4 L. 2.

mc06se cFMSr i-vi - Ed W. Clark High School

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SECTION 3 Date CHAPTER 11 REVIEW
Gases Class SHORT ANSWER Answer the following questions in the space provided c c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol (b) divided by the mass of 1 mol nRT (c) multiplied by 224 L (d) divided by 224 L For the expression $V =$ (a) increasing P

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changed into petroleum and natural

gases. The oil and gas moved through
porous rocks or along cracks until it was
trapped under impermeable rocks. ...

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CHAPTER 11 REVIEW Molecular
Composition of Gases SECTION 11-1
SHORT ANSWER Answer the following
questions. Point #1 Gas Volume
Relationships Modern Text Pg 378 -381.

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